**Chapter-20**

**1. A spherical balloon of volume 4.00 x 103 cm3 contains helium at a pressure of 1.20 x 105 Pa. How many moles of helium are in the balloon if the average kinetic energy of the helium atoms is 3.60 x 10-22 J?**

**Solution:**

**2. Calculate the mass of an atom of (a) helium, (b) iron, and (c) lead. Give your answers in kilograms. The atomic masses of these atoms are 4.00 u, 55.9 u, and 207 u, respectively.**

**Solution:**

**3. A 7.00-L vessel contains 3.50 moles of gas at a pressure of 1.60 x 106 Pa. Find (a) the temperature of the gas and (b) the average kinetic energy of the gas molecules in the vessel. (c) What additional information would you need if you were asked to find the average speed of the gas molecules?**

**Solution:**

**4. At what temperature would the average speed of helium atoms equal (a) the escape speed from the Earth, 1.12 x 104 m/s, and (b) the escape speed from the Moon, 2.37 x 103 m/s? Note: The mass of a helium atom is 6.64 x 10-27 kg.**

**Solution:**

**5.** **In a period of 1.00 s, 5.00 3 1023 nitrogen molecules strike a wall with an area of 8.00 cm2. Assume the molecules move with a speed of 300 m/s and strike the wall head-on in elastic collisions. What is the pressure exerted on the wall? Note:The mass of one N2 molecule is 4.65 3 10226 kg.**

**Solution:**